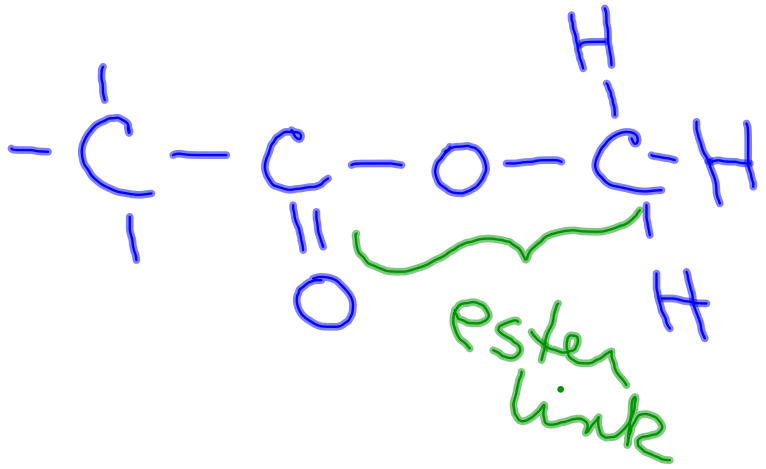
$$Cr\overset{-2}{O}_4^{\,2-}$$

$$-8 + x = -2$$

$$x = +6$$

$$CH_3COOH + CH_3OH \longrightarrow H_2O + CH_3COOCH_3$$



$$2H_2O + Cl_2 \xrightarrow{\text{oxid.}} 2\overset{+1}{C}lO^- + \overset{4}{\cancel{2}}H^+ + 2e^-$$

$$2e^- + Cl_2 \xrightarrow[\text{Red.}]{} 2Cl^-$$

$$5e^- + 8H^+ + \overset{+7}{Mn}\overset{-2}{O_4}^-_{red} \rightarrow Mn^{2+} + 4H_2O$$

$$H_2O + \overset{+3}{N}\overset{-2}{O_2}^-_{ox.} \rightarrow \overset{+5}{N}\overset{-2}{O_3}^- + 2H^+ + 2e^-$$